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This reaction follows the equation: $N_2 + 3H_2 \rightarrow 2NH_3$ $\Delta H^\circ = -92 \text{ kJ}$

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mol $S_0 = 192 \text{ J mol}^{-1} \text{ K}^{-1}$ To find the Gibbs free energy of formation at room temperature, recall that $G_0 = H_0 - T S_0(2) = 92 \text{ kJ mol}^{-1} + (298 \text{ K}) (0.192 \text{ kJ mol}^{-1} \text{ K}^{-1}) = 35 \text{ kJ mol}^{-1}$ Alternatively, one can also find the temperature for which $G = 0$, $T = H_0 / S_0$.

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